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gases, liquids, or solids. 2) Explain why a suspension is considered a heterogeneous mixture. The particles are so large that they settle out unless the mixture is constantly stirred or agitated.

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The following are statements about the dissolving process. Explain each one at the molecular level. a. Increasing the pressure of a solute gas above a liquid solution increases the solubility of the gas in the liquid.

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in the space provided. 1. Match the type of mixture on the left to its representative particle diameter on the right. c solutions (a) larger than 1000 nm a suspensions (b) 1 nm to 1000 nm b colloids (c)

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right. c solutions (a) larger than 1000 nm a suspensions (b) 1 nm to 1000 nm b colloids (c) smaller than 1 nm 2. Identify the solvent in each of the following examples:

*mc06se cFMSr i-vi*

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*Chapter 12, Problem Review\_and\_Discussion 13 : 5. How does ...*

Modern Chemistry Chapter 12 Solutions Review Answers Modern Chemistry Chapter 12-Solutions Section 1- Types of Mixtures Solutions are homogeneous mixtures of two or more substances in a single phase. Soluble describes a substance as capable of being dissolved. Solvent is the dissolving medium in a solution. Chapter 12 Review Solutions Section 3 Answers

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