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Properties Molality

and Colligative

Properties

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Solutions 14.4

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What are

Solutions? Solute,

Solvent and

Solution |

Chemistry

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Elevation Boiling

Point Elevation and

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Depression

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Types of Solution |

What is a solution?

Chemistry

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Diagrams of Water

\u0026 CO<sub>2</sub>

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\u0026 Critical

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~~PROPERTIES |~~

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~~2019 |~~

~~Solutions(Osmotic  
Pressure)~~

~~Chemistry (L-13) |~~

~~by Arvind Arora~~

~~SOLUTION /PART 5/~~

~~CLASS 12/~~

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~~PROPERTY/REVERSE~~

~~OSMOSIS/VANT~~

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Two solutions with the same osmotic pressure (same concentration of solutes) are said to be isotonic. Why must intravenous fluids be isotonic

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with your blood?

Osmotic pressure is a colligative property of a solution. That is, its magnitude depends on the concentration of dissolved particles but does not depend on the nature of the dissolved particles. Interestingly,

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osmotic pressure  
( $\Pi$ ) can be  
calculated using an  
equation that is  
very similar to the  
ideal gas equation:

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Properties

(Worksheet) -

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Worksheet #4:

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Properties \*Starred  
problems are  
especially good  
/challenging

practice FP

Depression/BP

Elevation: 12.77,

12.79, 12.85,

12.87, 12.103\*,

12.131 Vapor

Pressure: 12.71,

12.73, 12.75,

12.91, 12.101\*,

12.107\*, 12.109\*,

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12.111\*, 12.123\*

Osmotic Pressure:

12.81, 12.83,

12.89, 12.121\* 1.

Answers File

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TIONS AND

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PROPERTIES SET A:

1. Find the molarity

of all ions in a

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Worksheet  
contains 0.165 moles of aluminum chloride in 820. ml

Answers File  
solution. Answer:

Type  
[Al<sup>3+</sup>] = 0.201 M ,  
(Cl<sup>-</sup>) = 0.603M. 2.

Find the molarity of each ion present after mixing 27 ml of 0.25 M HNO<sub>3</sub> with 36 ml of 0.42 M Ca(NO<sub>3</sub>)<sub>2</sub>

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Worksheet Answers

1)  $100.11\text{ }^{\circ}\text{C}$  2)

$100.12\text{ }^{\circ}\text{C}$  3)

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boiling point =  
102.8 °C. freezing  
point = -10 °C 4)

Use the following  
info:  $K_f = 4.68$

°C/m. Normal F.P =  
-63.5 °C 5) NaCl,

HCl, CaCl<sub>2</sub>,

C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>,

HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> 6) Use

the following info:

$K_b = 5.02$  °C/m

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Worksheet II - blogs

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PROPERTIES SET A:

1. Find the molarity

of all ions in a

solution that

contains 0.165

moles of aluminum

chloride in 820. ml

solution. Answer:

[Al<sup>3+</sup>] = 0.201 M ,

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$[Cl^-] = 0.603 M$ .

Find the molarity of each ion present

after mixing 27 ml

of 0.25 M  $HNO_3$

with 36 ml of 0.42

M  $Ca(NO_3)_2$

(Note: There is no reaction taking place.)

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properties are

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molality. As a

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definition of



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molality: Because the vapour pressure of a solution with a nonvolatile solute is depressed compared to that of the pure solvent, it requires a higher temperature for the solution's vapour pressure to reach 1.00 atm (760 torr).

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Solutions  
Properties of  
Worksheet  
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Introductory ...

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properties are  
properties of  
solutions, that  
depend on the  
concentration of  
the dissolved  
particles  
(molecules or ions),

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Properties Of  
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but not on the identity of those particles. They often affect solvent properties like boiling and melting point, or the vapor pressure above a fluid. There are four colligative properties we will look at, which are:

13.4: Colligative

*Page 35/52*

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worksheet,

students determine

the boiling points

and melting points

of solutions.

Students calculate

the effective

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Properties of a solute.

This worksheet has five problems to solve.

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discussed,

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Solutions have different properties than either the solutes or the solvent used to make the solution. Those properties can be divided into two main groups--colligative and non-colligative properties. Colligative properties depend



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only on the number  
of dissolved  
particles in solution  
and not on their  
identity.

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Solutions:

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Properties

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properties are the

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properties of a solution as a whole and depend on the concentration. The colligative properties include freezing point depression, boiling point elevation, vapor pressure lowering and osmotic pressure. An overview of the colligative

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best of your ability.

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counted against  
you. When you  
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Raoult's Law ...

There are a few  
solution properties,  
however, that  
depend only upon  
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concentration of solute species, regardless of their identities. These colligative properties include vapor pressure lowering, boiling point elevation, freezing point depression, and osmotic pressure.



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